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change. Gibbs free energy and spontaneity. Gibbs free energy example. More rigorous Gibbs free energy / spontaneity relationship.

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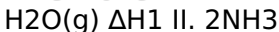
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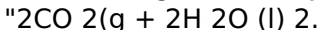
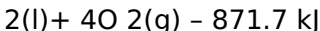
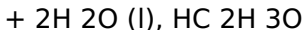
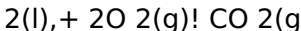
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words ΔH is positive? I.



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ANSWERS, 1. $\text{HC}_2\text{H}_3\text{O}_2(\text{l}) + 2\text{O}_2(\text{g}) \rightarrow$



The enthalpy of

reaction is $\Delta H^\circ_{\text{prod}}$

$-\Delta H^\circ_{\text{react}}$. The

reactant is at a higher

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potential than the product. When the reaction occurs, heat is released, stabilizing the system and the system moves to a lower energy potential.

3.

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Thermochemistry practice problems 1)
How can energy be transferred to or from a system? A) Energy can

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only be transferred as potential energy being converted to kinetic energy. B) Energy can be transferred only as heat. C) Energy can be transferred only as work. D) Energy can be transferred as heat and/or work.

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6. When 26.7 g of H₂S

was burned in excess

oxygen, 406 kJ was

released. What is H for

the following equation?

$2 \text{H}_2\text{S}(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow$

$\text{SO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g});$

H = ??? Answer: -1040

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kJ Reasoning: $2 \times 26.7\text{g}$

$0.78345 \text{ mol H}_2\text{S}$ 34.08

$\text{g/mol H}_2\text{S}$, so 406 kJ

was released when

0.78345 moles of H_2S

reacted. 2×2

Answers,

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Name_____ MULTIPLE

CHOICE. Choose the

one alternative that

best completes the

statement or answers

the question. 1) A

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A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.

A) endothermic,
positive

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Temperature

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study of changes in
energy (heat)

associated ... notice
final answer in

problems above should
be 3 sig fig 2.09×10^4 J
or 20.9 kJ .

Thermochem 9

Calorimeter device to
measure changes in

heat Bomb (metal
chamber) Calorimeter
shown below ...

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Answers

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Example #4: Lead has a melting point of $327.5\text{ }^{\circ}\text{C}$, its specific heat is $0.128\text{ J/g}\cdot^{\circ}\text{C}$, and its molar enthalpy of fusion is 4.80 kJ/mol . How much heat, in kilojoules, will be required to heat a 500.0 g sample of lead from $23.0\text{ }^{\circ}\text{C}$ to its melting point and then

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melt it?

**ChemTeam:
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Problems - two
equations needed**

For each of the following questions or statements, select the most appropriate response and click its letter:

**Quiz #3-3 PRACTICE:
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Mr. Carman's Blog**

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ANSWER KEY 1.

Calculate the standard enthalpy of reaction for the process $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$ $\Delta H =$

? 2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}

$\text{C}_2\text{H}_5\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) \rightarrow$

$2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$

$\Delta H = ?$

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consider 2 metals a

and b each having a

mass of 100 g and an

initial temperature of

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answers ...

Chapter 17

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Answers 1. What will

be sign for q and W if

an isolated system

absorb energy from the

surrounding and does

work for expansion. 2.

The amount of work

done in joules by the

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Answers
system in expanding from 1.50L to 2.3L against a constant atmospheric pressure of about 1.3atm.

1. 2 3. -

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Tags: Question 3 .

SURVEY Q.

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study of ΔH changes
during chemical and
physical reactions.
answer choices

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Thermochemistry 6
Exercise 5 Constant-
Pressure Calorimetry
When 1.00 L of 1.00 M
 $\text{Ba}(\text{NO}_3)_2$ solution at
 25.0°C is mixed with
1.00 L of 1.00 M
 Na_2SO_4 solution at
 25°C in a calorimeter,

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the white solid BaSO_4 forms and the temperature of the mixture increases to 28.1°C . Assuming that the calorimeter absorbs only a negligible quantity of heat, that the specific heat

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