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Note as well that since the system is isolated, no heat has been exchanged with the surroundings ($q = 0$). The first law of thermodynamics confirms that there has been no change in the system's internal energy as a result of this process. \[

$$\begin{aligned} \Delta U &= q + w \tag{First Law of Thermodynamics} \\ &= 0 + 0 = 0 \end{aligned}$$

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