

Chapter 4 Atomic Structure Test A Answers

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CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure Matching. A. Bohr B. Democritus C. Rutherford D. Dalton E. Thomson F. Schrodinger ____ 1. Greek thinker; called nature's basic particle an atom, based on the Greek word "atomos" which means "indivisible". Did not have evidence that atoms existed. ____ 2.

CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure

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a student of Thompson, who tested the current theory of the atomic structure, using alpha particles (positively charged) which are directed at a sheet of gold foil, according to thompson's theory the particles should have passed through the gold with only slight deflection due to the positive charge in the foil, but he proved this wrong because majority of the particles passed through the gold ...

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Chapter 4 Test - Chemistry Discuss the difference between atomic number, mass number, and average atomic mass. 4. Discuss the significance of Rutherford's findings regarding the atom's structure. 5. An element that is very reactive gas is most likely a member of the. A. Alkali metals. B.

Chapter 4 Test - Chemistry - ProProfs Quiz

Section 4.1 Defining the Atom. OBJECTIVES: Identify what instrument is used to observe individual atoms. Section 4.1 Defining the Atom. The Greek philosopher Democritus (460 B.C.- 370 B.C.) was among the first to suggest the existence of atoms (from the Greek word "atomos") He believed that atoms were indivisible and.

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ICSE Class7 Chemistry Chapter 4 Atomic Structure Part-2

Chapter 4 "Atomic Structure" Tools. Copy this to my account; Email to a friend; ... Help; Use these activities to help your study the vocabulary terms from this chapter. A B; atomic mass unit: defined as the mass of a carbon -12 atom: nucleus: central core of an atom, which contains most of the atom's mass: group: a vertical column of elements ...

Quia - Chapter 4 "Atomic Structure"

Chapter 4 - Atomic Structure - Standardized Test Prep - Page 125: 4. Answer. C. Work Step by Step. The top number will be the mass number and the bottom number would be the atomic number. The mass number minus the atomic number will equal the number of neutrons. The atomic number is equivalent to the number of protons and therefore if can be determined that there are 16 protons.

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Atomic Structure - Unit Test Paper 4. Q.1. Select the correct answer from the answers in brackets to complete each sentence. Question 1. An element has electronic configuration 2, 8, 1 and 12 neutrons. Its mass no. is ____ [11 / 23 / 12] Answer: An element has electronic configuration 2, 8, 1 and 12 neutrons. Its mass no. is 23. Question 2.

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4 Atomic Structure • Chapter Test A. Matching Match each description in Column B with the correct term in Column A. Column A Column B 1. proton 4·2 a. the smallest particle of an element that retains the properties of that element 2. atom 4·} b. the weighted average of the masses of the isotopes of 3. mass number 4·5 an element 4. atomic mass unit 4·5

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Chemistry Atomic Structure Online Quiz Test MCQs. 1. The nature of positive rays depends on: Nature of electrode. Nature of resident gas. All of the above. Nature of discharge. Question 1 of 15. 2. Alpha rays consists of: Protons. Hydrogen nucleus. Helium nucleus. Neutrons. Question 2 of 15. 3. X-rays have same nature as: ...

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